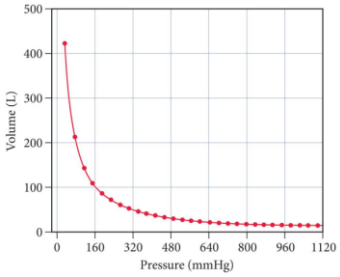
N26 – Ideal Gases and Laws

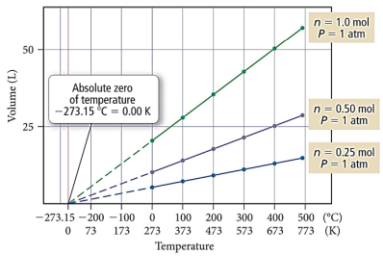
Ideal Gas Law
$PV = nRT$

Molar Mass Kitty
$M = \frac{DRT}{P}$

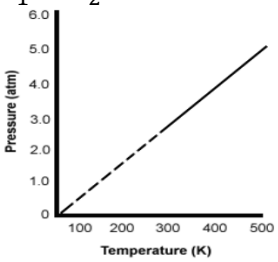
Boyle's Law
$P_1V_1 = P_2V_2$



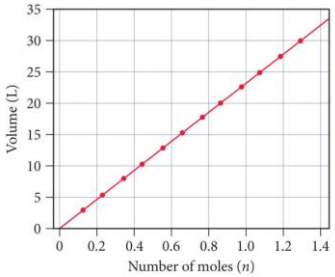
Charles's Law
$\frac{V_1}{T_1} = \frac{V_2}{T_2}$



Gay-Lussac's Law
$\frac{P}{T_1} = \frac{P_2}{T_2}$



Avogadro's Law
$\frac{V_1}{n_1} = \frac{V_2}{n_2}$



Dalton's Partial Pressure
$P_{total} = P_1 + P_2 + \cdots$

Mole Fraction
$X_a = \frac{n_a}{n_{total}} \qquad P_a = X_a P_{total}$

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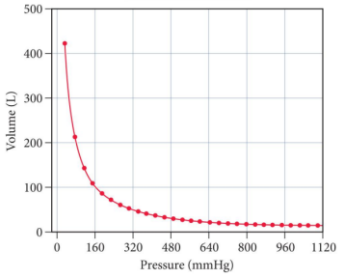

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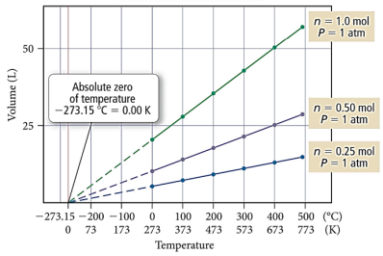

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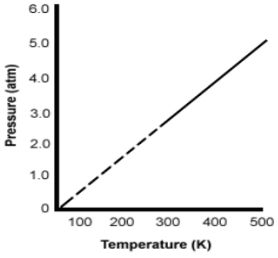

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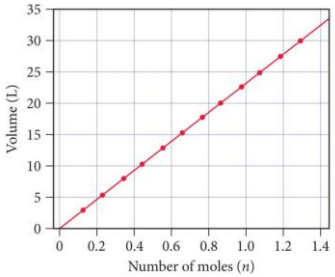

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